

Method of concentrating and removing hydrogen sulfide from a contaminated gas mixture

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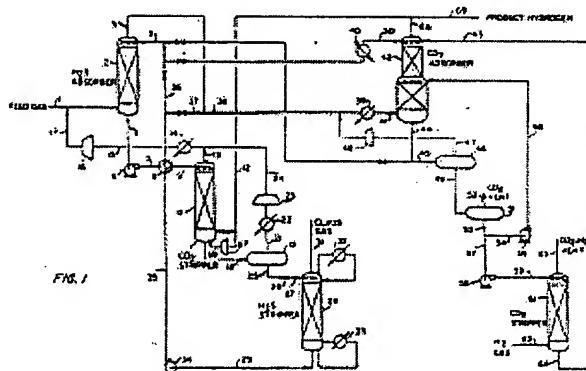
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Abstract of EP0033029

Method for treating a contaminated gas such as a contaminated natural gas or a synthesis product gas from the partial oxidation of petroleum or coal to concentrate and remove hydrogen sulfide at a reduced energy and capital requirement.

Hydrogen sulfide in the contaminated stream is absorbed to form a hydrogen sulfide rich solvent which includes co-absorbed carbon dioxide, which is then removed to prevent dilution of the concentrated hydrogen sulfide formed when the solvent is regenerated. The solvent leaving the hydrogen sulfide absorber (2) is heated before entering a high pressure stripper column (11) wherein the co-absorbed carbon dioxide is stripped from the solvent, preferably by stripping with nitrogen or a portion of the product gas, and the carbon dioxide is recycled to the hydrogen sulfide absorber. The CO₂ stripper is operated at a pressure not substantially below the pressure of the hydrogen sulfide absorber, eliminating the need for large compressors required in conventional methods to recycle the carbon dioxide.



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Method of concentrating and removing hydrogen sulfide from a contaminated gas mixture

Description of corresponding document: **US4242108**

BACKGROUND OF THE INVENTION

This invention relates to a method for concentrating and removing hydrogen sulfide from a contaminated gas mixture containing hydrogen sulfide and carbon dioxide, such as may be found in contaminated natural gas or the contaminated product gas from coal gasification or from the partial oxidation of petroleum.

Numerous processes for removing hydrogen sulfide and carbon dioxide from gaseous mixtures involving physical absorption are well known. Kutsher et al, U.S. Pat. No. 3,362,133, discloses a process for removing hydrogen sulfide and carbon dioxide by physical absorption in the dialkylether of a polyalkylene glycol and then flashing the solvent at a pressure substantially lower than the absorber to release any co-absorbed CO₂, which is recycled to the absorber. The flashing step separates CO₂ from the solvent, but any recycle can be effected only by compressing the CO₂ before it can be recycled into the absorber which is at a substantially higher pressure than the flashed CO₂. Tjoa et al., U.S. Pat. No. 3,767,766, uses a process similar to Kutsher and shows passing the solvent through an expansion valve to flash and vaporize dissolved CO₂ and/or hydrocarbons which may be recycled through a compressor to the absorber.

Hill, U.S. Pat. No. 3,989,811, shows a sulfur recovery process in which an absorbent containing H₂S and CO₂ is heated and then stripped of H₂S and CO₂ in a low pressure stripping column to regenerate the solvent. The stripped vapors are passed to a second H₂S absorber to separate CO₂ from the H₂S, and the CO₂ is vented.

Child et al, U.S. Pat. No. 4,520,176, shows a process for removing acid gases including CO₂, H₂S, and COS, from a raw synthesis gas feed stream by physical absorption. The hydrogen sulfide rich absorbent is stripped in low pressure columns using a portion of the H₂-rich product gas stream as an optional stripping gas to regenerate purified solvent. The gas stream leaving the stripper is introduced into underground reservoirs to facilitate petroleum recovery.

The conventional processes for separating hydrogen sulfide and carbon dioxide from gaseous mixtures, as represented by the foregoing, are limited by the low pressure techniques of separating any co-absorbed carbon dioxide from the solvent rich in absorbed hydrogen sulfide, such as by expansion and flashing or by low pressure stripping columns. Using the conventional methods, large and costly compressors are required to effect a recycle back to an H₂S absorber which necessarily can be operated only at a pressure substantially higher than the pressure in the flashing separation vessel.

An object of this invention is to overcome the limitations of the conventional processes and provide for an efficient separation of any co-absorbed carbon dioxide in the hydrogen sulfide rich solvent leaving the H₂S absorber. It is a further object of this invention to provide for an efficient separation and removal of said co-absorbed CO₂ without the need for large and expensive compressors to recycle any stripped CO₂ into the high pressure H₂S absorber. It is a further object of this invention to provide for the use of the portion of the product gas from the contemplated gas cleanup as the stripping gas in the CO₂ stripper for removing any co-absorbed CO₂ from the hydrogen sulfide rich solvent. It is yet a further object of this invention to adjust the portion of product gas used in the CO₂ stripper so as to provide for a H₂S concentration higher than conventional methods in the product Claus gas for use in a sulfur recovery unit.

SUMMARY OF THE INVENTION

According to our invention a contaminated natural gas or synthesis gas mixture having components of hydrogen sulfide and carbon dioxide, is passed through a hydrogen sulfide absorber to remove the hydrogen sulfide and which also removes a portion of the carbon dioxide. Methane, hydrogen, and/or a mixture of hydrogen and carbon monoxide and the remaining carbon dioxide are withdrawn from the top of the H₂S absorber, and CO₂ is removed to form a product gas. The hydrogen sulfide rich solvent leaving

the absorber is heated and then passed to a high pressure CO₂ stripping column wherein any co-absorbed CO₂ is stripped from the solvent using any stripping gas such as nitrogen from an air separation plant, or preferably an adjustable portion of the product gas. The stripped CO₂ then is recycled to the H₂ S absorber without the need for large compressors to bring the CO₂ back up to the pressure necessary for admission into the H₂ S absorber. The hydrogen sulfide rich solvent leaving the CO₂ stripper then may be regenerated to provide a high H₂ S concentration Claus gas and a purified solvent.

BRIEF DESCRIPTION OF THE DRAWING

In the drawing:

FIG. 1 is a simplified schematic diagram of a process system for a practical operation in the practice of the novel method according to the invention.

DETAILED DESCRIPTION

Reference is directed to FIG. 1 wherein the novel method of the invention is illustrated by a simplified block flow diagram. A feed gas such as a contaminated natural gas containing methane or a contaminated synthesis gas containing hydrogen such as from a shifted product gas mixture from a coal gasification process or a partial oxidation of petroleum, at a pressure between 400 and 2,500 psia, and preferably between 600 and 1,000 psia, and also more particularly containing components of carbon dioxide, and hydrogen sulfide, and which may also contain carbon monoxide, argon, nitrogen, carbonyl sulfide, and water, is passed in stream 1 to an H₂ S absorber 2. The feed gas is contacted with a liquid absorbent, preferably having a higher selectivity for absorbing H₂ S over CO₂, which enters the top of H₂ S absorber in line 3. Hydrogen sulfide is absorbed and, along with any co-absorbed CO₂, leaves the H₂ S absorber in the liquid solvent in line 4. Vapors which are not absorbed in the liquid solvent, such as carbon dioxide; methane, particularly in the case when the feed gas is a contaminated natural gas; and hydrogen or a mixture of hydrogen and carbon monoxide, particularly in the case when the feed gas is a contaminated synthesis gas; leave the H₂ S absorber in line 5.

The solvent leaving the bottom of the H₂ S absorber in line 4 is pumped at 6 and passed in line 7 to be heated in exchanger 8 before entering in line 9 the top of a high pressure CO₂ stripper column 1 at a pressure not substantially below the pressure in H₂ S absorber 2. A stripping gas, such as nitrogen or a portion of the product gas from downstream processing, which will be predominantly methane when the feed gas is natural gas and which will be hydrogen or a mixture of hydrogen and carbon monoxide when the feed is a synthesis gas, is admitted in line 12 to the bottom of CO₂ stripper 11. The stripping gas in line 12 can be N₂ from an air separation plant such as in the case of the H₂ product gas being used as a feed to an ammonia plant. CO₂ rich gas is stripped from the liquid solvent and leaves the CO₂ stripper as overhead in line 13 which must be recycled to H₂ S absorber 2 to remove any H₂ S contained in this gas stream. The CO₂ rich gas in line 13, at a pressure not substantially below that pressure used in line 1, is cooled at 14. CO₂ rich gas leaving cooler 14 in line 15 is compressed slightly at 16 to make up for minor pressure losses and is recycled in line 17 to join the feed gas in line 1.

Hydrogen sulfide rich absorbent leaves the bottom of the CO₂ stripper 11 in line 18 and is flashed at 19 whereupon coabsorbed gas, such as hydrogen, methane, carbon dioxide, and carbon monoxide, is removed for recovery in line 21, is cooled in exchanger 22, recompressed at 23, and recycled in line 24 by admixing with the CO₂ recycle in line 13. The hydrogen sulfide rich solvent leaves the flash vessel as line 25 and is expanded in a Joule-Thompson expansion valve 26 and passed in line 27 to H₂ S stripper 28 wherein the solvent is regenerated and leaves the bottom as purified solvent in line 29. Concentrated H₂ S leaves the top of the H₂ S stripper in line 31 and may be delivered as a Claus gas having a high H₂ S concentration to a sulfur recovery unit. Steam is generated in reboiler 33 to strip the H₂ S, and condenser 32 condenses the water at the top of stripper 28.

Lean solvent leaving the bottom of the H₂ S stripper in line 29 is passed through pump 34 and passes in line 35 in indirect heat exchange with solvent in line 7 in heat exchanger 8. After losing heat to solvent in line 7, the purified solvent can either join line 3 in line 36 or preferably as line 37 joins the vapors in line 5 to form a vapor-liquid stream in line 38 or further can optionally be used after refrigeration in exchanger 40 as absorber feed in line 50 to the top of CO₂ absorber 42.

Solvent and the gases from the overhead of the H₂ S absorber in line 38 are cooled in exchanger 39 and enter the bottom of a CO₂ absorber 42 through line 41. A pure liquid absorbent enters the top of CO₂ absorber 42 in line 43. CO₂ is absorbed, and a CO₂ rich liquid absorbent leaves the bottom of the CO₂ absorber in line 44. A portion of the CO₂ rich liquid absorbent in line 44 is passed in line 3 to the top of H₂ S absorber 2. The remaining CO₂ rich solvent in line 45 is flashed in flash vessel 46. Co-absorbed gas which is flashed in vessel 46 is conveyed in line 47 through compressor 48 and admixed with the solvent and vapors in line 38.

CO₂ rich liquid absorbent leaves the bottom of the flash vessel 46 in line 49 and is flashed in flash vessel 51. Carbon dioxide gas leaves the flash vessel 51 through line 52 and may be vented. Liquid absorbent leaves the flash vessel 51 through line 53 and is split so that a portion can be passed in line 54 through pump 55 to CO₂ absorber 24 and enters at an intermediate portion of the absorber.

A second portion of the liquid absorbent leaving flash vessel 51 in line 53 is passed in line 57 through pump 58 to the top of a CO₂ stripper 61 in line 59. Nitrogen gas which can be supplied from an air separation plant is passed in line 62 to the bottom of CO₂ stripper 61. Waste nitrogen and carbon dioxide leaves the top of CO₂ stripper 61 in line 63 and can be vented. Lean liquid absorbent leaves the bottom of CO₂ stripper 61 in line 64 and is passed through pump 65 before passing to and entering the top of CO₂ absorber 42 in line 43.

Product gas leaves the top of CO₂ absorber 42 in line 66. A portion of the product gas in line 66 is passed in line 12 to the bottom of CO₂ stripper 11 and serves as the stripping gas. The gas in line 12 optionally can be slightly compressed at 67 to increase the stripping gas pressure in line 68 by 10-30 psi before entering CO₂ stripper 11, which slight compression will eliminate the need for compressor 16. Hydrogen product gas leaves the process in line 69.

The portion of the product gas used as a stripping gas in the high pressure CO₂ stripping column can be adjusted to give a higher concentration of hydrogen sulfide in the Claus gas particularly in comparison with conventional methods for concentrating and removing H₂ S from contaminated natural gas or synthesis gas. Increasing product gas volumes used as a stripping gas will increase the concentration of hydrogen sulfide from the H₂ S stripper and provide for a more efficient sulfur recovery.

The CO₂ stripper column is operated at a pressure not substantially below the pressure of the hydrogen sulfide absorber, thereby eliminating requirements for large compressors to recycle the carbon dioxide back to the high pressure H₂ S absorber. For example, operation of the CO₂ stripper column 11 at a pressure about 10 to 30 psi below the pressure in hydrogen sulfide absorber 2 is suitable.

Other advantages of the novel method according to this invention may be seen from an inspection of the following example.

EXAMPLE

The following example is offered for a better understanding of the present invention, but the invention is not to be construed as limited thereto. This example is applied to a plant which processes a feed gas comprised of:

<tb> _____
<tb> Mole %
<tb> _____
<tb> Carbon Monoxide 0.40
<tb> Hydrogen 62.62
<tb> Carbon Dioxide 35.46
<tb> Methane 0.36
<tb> Argon 0.34
<tb> Nitrogen 0.07
<tb> Hydrogen Sulfide 0.55
<tb> Carbonyl Sulfide 0.003
<tb> Water 0.20
<tb> _____

The feed gas is introduced into the H₂ S absorber 2 at a pressure of about 654 psia and at about 104 DEG F. The fresh feed flow rate is designed at about 13,120 lb. moles gas per hour, to which is added

1040 lb. moles gas per hour of recycle gas in line 17. The solvent stream from the bottom of the CO₂ absorber 42 is introduced at about 40 DEG F. into absorber 2 by line 3 at a rate to provide 1350 lb. moles/hour of solvent comprised of dimethyl ether of polyethylene glycol. In addition to the organic solvent, the stream also contains carbon dioxide (35.1 mole %), water (17.7 mole %), and various other components in small concentrations. As the solvent flows down absorber 2 most of the H₂ S in the feed gas is absorbed along with other co-absorbed species resulting in an essentially H₂ S free overhead in line 5. The rich solvent leaving absorber 2 in line 4 at a flow rate of 3010 lb. moles/hour is comprised of:

<tb> _____
<tb> Mole %
<tb> _____
<tb>Carbon Monoxide 0.01
<tb>Hydrogen 0.60
<tb>Carbon Dioxide 31.20
<tb>Methane 0.02
<tb>Argon --
<tb>Nitrogen --
<tb>Hydrogen Sulfide 4.51
<tb>Carbonyl Sulfide 0.008
<tb>Water 18.78
<tb>Dimethyl Ether of 44.88
<tb>polyethylene glycol
<tb> _____

The rich solvent in line 4 at a temperature of about 75 DEG F. and pressure of 654 psia is increased in pressure to about 660 psia in pump 6 before passing in line 7 into heat exchanger 8. The rich solvent is heated in exchanger 8 to between 265 DEG F. and 300 DEG F. by indirect heat exchange with line 35 before passing into stripper 11 in line 9. A maximum of 1% of the product hydrogen in line 66 is by-passed in line 12 to serve as a strip gas in absorber 11. The gas in line 12, at about 15.5 DEG F. and 639 psia, also contains carbon monoxide (0.62 mole %), methane (0.55 mole %), argon (0.53 mole %), nitrogen (0.12 mole %), and small concentrations of other compounds. The total flow of stream 12 into absorber 11 is a maximum of 85 lb. moles/hour. As the hydrogen strip gas passes up absorber 11 the carbon dioxide is stripped from the solvent along with small amounts of other absorbed species. The compositions of the vapor leaving stripper 11 in line 13 and the composition of liquid stream 18 at the bottom of stripper 11 are comprised of:

<tb> _____
<tb> Stream 13 Stream 18
<tb> (mole %) (mole %)
<tb> _____
<tb>Carbon Monoxide 0.06 0.01
<tb>Hydrogen 8.15 1.06
<tb>Carbon Dioxide 83.58 6.45
<tb>Methane 0.07 0.02
<tb>Argon 0.04 --
<tb>Nitrogen 0.01 --
<tb>Hydrogen Sulfide
<tb> 5.64 3.81
<tb>Carbonyl Sulfide
<tb> 0.01 --
<tb>Water 2.40 25.38
<tb>Dimethyl Ether of
<tb> -- 63.27
<tb>Polyethylene Glycol
<tb> _____

The rich solvent in line 18 at about 262 DEG F. and 639 psia now contains considerably less carbon dioxide than the solvent leaving absorber 2 in line 4. The liquid, in line 18, at a flow rate of about 2134 lb. moles/hour, is then flashed down to a pressure in the range of 200-300 psia in flash drum 19 to desorb hydrogen from the solvent. The liquid leaving flash drum 19 at about 260 DEG F. passes through a Joule-Thomson valve 26 before entering H₂ S stripper 28 where steam is used to regenerate the solvent and to produce a Claus gas in line 31 containing at least 40 mole % hydrogen sulfide. The lean solvent in line 29 is increased in pressure through pump 34 and after heat exchange in exchanger 8 is mixed with line 5.

The vapor leaving flash drum 19 in line 21 at about 260 DEG F. and at a flow rate of about 83 lb. moles/hour is comprised of:

<tb> _____
<tb> Moles %
<tb> _____
<tb>Carbon Monoxide 0.23
<tb>Hydrogen 23.13
<tb>Carbon Dioxide 59.98
<tb>Methane 0.25
<tb>Argon 0.03
<tb>Nitrogen 0.01
<tb>Hydrogen Sulfide 10.82
<tb>Carbonyl Sulfide 0.02
<tb>Water 5.47
<tb>Dimethyl Ether 0.06
<tb>Polyethylene Glycol
<tb> _____

The vapor in line 21 is cooled to about 105 DEG F. in heat exchanger 22 with cooling water before passing through compressor 23 to give a pressure in line 24 of 635 psia. The vapor in line 24 is mixed with the vapor in line 13, and this mixture is cooled in heat exchanger 14 to about 105 DEG F. in line 15 before passing through compressor 16 to give the same pressure as the feed gas in line 1.

An advantage of using the novel method according to our invention in the process described in this example is that the compressors used in recompressing the desorbed CO₂ are much smaller than those used in conventional methods of concentrating the Claus gas. Because CO₂ stripper 11 is maintained at a relatively low pressure differential compared to H₂ S absorber 2 (19 psia), the bulk of the CO₂ desorbed from the solvent in line 13 requires a small pressure differential of compression (24 psia) in compressor 16 to reach the inlet feed pressure. Thus about 92% of the total gas in line 17 is compressed over a relatively small pressure differential of compression compared to conventional flash recompression schemes where most of the desorbed gas would be compressed over pressure differentials of compression up to 550 psia.

A further advantage of using the novel method described in this example according to our invention is that a flexibility is provided to increase the H₂ S concentration in the Claus gas with minimal incremental costs. By adjusting the H₂ flow in line 12, more CO₂ can be desorbed from the solvent in stripper 11, and thus a higher H₂ S Claus gas concentration is achieved in line 31.

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Method of concentrating and removing hydrogen sulfide from a contaminated gas mixture

Claims of corresponding document: **US4242108**

What is claimed is:

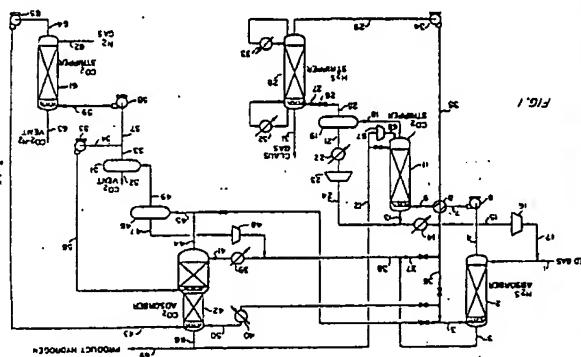
1. A method of removing hydrogen sulfide from a contaminated gas mixture comprising hydrogen sulfide, carbon dioxide, and a gas selected from the group consisting of hydrogen, methane, and a mixture of hydrogen and carbon monoxide, which method comprises: (a) contacting said contaminated gas mixture with a liquid absorbent in a first liquid-gas contacting zone to absorb said hydrogen sulfide and form a hydrogen sulfide rich liquid absorbent and an overhead gas, and thereby absorbing a portion of said carbon dioxide, (b) removing carbon dioxide from said overhead gas leaving said first contacting zone in a carbon dioxide absorber to form a product gas, (c) heating said hydrogen sulfide rich liquid absorbent leaving said first contacting zone to form a heated hydrogen sulfide rich liquid absorbent, (d) stripping carbon dioxide from said heated hydrogen sulfide rich liquid absorbent with a stripping gas in a second liquid-gas contacting zone to form a carbon dioxide rich gas and an effluent hydrogen sulfide rich liquid absorbent stream containing stripping gas, (e) cooling, compressing, and recycling said carbon dioxide rich gas to said first contacting zone, (f) separating stripping gas from said hydrogen sulfide rich liquid absorbent stream leaving said second contacting zone in a first liquid-gas separating zone to form a separated stripping gas and a stripped hydrogen sulfide rich liquid absorbent, and admixing said separated stripping gas with said carbon dioxide rich gas leaving said second contacting zone, (g) separating hydrogen sulfide from said stripped hydrogen sulfide rich liquid absorbent leaving said first separation zone in a second liquid-gas separation zone to form hydrogen sulfide gas and purified liquid adsorbent, and (h) recycling said purified liquid adsorbent from said second liquid-gas separation zone to said first liquid-gas contacting zone with from 0 to 100 percent of said purified liquid adsorbent being recycled passing first through said carbon dioxide absorber.
2. A method according to claim 1 wherein said stripping in said second contacting zone is effected at a stripping pressure not substantially below a pressure in said first contacting zone.
3. A method according to claim 2 wherein said stripping gas is selected from the group consisting of a portion of said product gas and nitrogen.
4. A method according to claim 3 wherein said liquid adsorbent has a higher solubility selectivity for hydrogen sulfide than for carbon dioxide.
5. A method according to claim 4 wherein said stripping pressure in said second contacting zone is in the range of 400-2,500 psia.
6. A method according to claim 4 wherein said stripping pressure in said second contacting zone is in the range of 600-1,000 psia.
7. A method according to claim 6 wherein said heating includes heating by indirect heat exchange communication with said purified liquid adsorbent leaving said second separation zone.
8. A method according to claim 7 wherein said removing carbon dioxide from said overhead gas leaving said first contacting zone comprises absorbing by said purified liquid adsorbent in a CO₂ absorber prior to said recycling of said purified liquid adsorbent to said first contacting zone.
9. A method according to claim 1 or 8 wherein said liquid adsorbent comprises dialkyl ether of polyethylene glycol.
10. A method according to claim 9 wherein said contaminated gas mixture comprises contaminated natural gas and wherein said product gas comprises methane.
11. A method according to claim 9 wherein said contaminated gas mixture comprises a contaminated synthesis gas from a partial oxidation of petroleum or coal, and wherein said product gas comprises a gas

selected from the group consisting of hydrogen and a mixture of hydrogen and carbon monoxide.

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0033029

11



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